

Single, multi-walled, functionalized and doped carbon nanotubes and composites thereof

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1. BACKGROUND OF THE INVENTION

Field of the Invention

The present invention relates to single walled and multi-walled carbon
10 nanotubes (CNTs), functionalized carbon nanotubes and carbon nanotube
composites with controlled properties, to a method for aerosol synthesis of single
walled and multi-walled carbon nanotubes, functionalized carbon nanotubes and
carbon nanotube composites with controlled properties from pre-made catalyst
15 particles and a carbon source in the presence of reagents and additives, to functional,
matrix and composite materials composed thereof and structures and devices
fabricated from the same in continuous or batch CNT reactors.

Description of related Art

Carbon nanotubes are of great interest since they exhibit unique and useful
20 chemical and physical properties related to, for instance, their morphology,
toughness, electrical and thermal conductivity and magnetic properties. Since their
discovery, CNTs have been the subject of intensive research and numerous patents,
scientific articles and books have been devoted to their synthesis, properties and
applications. Nanotubes were first observed during a direct-current arc discharge
25 between graphite electrodes in an argon environment by Iijima (Nature 1991, **354**,
56). The typical temperatures for carbon nanotube production by that method are
about 2000-3000 °C. Since then, various authors described alternative means of
carbon nanotubes production, which allowed increased production rate and
significantly decreased temperatures, e.g., [Jiao and Seraphin, *J. Phys. & Chem.*
30 *Solids* 2000, **61**, 1055; Hafner et al., *Chem. Phys. Lett.* 1998, **296**, 195]. For
instance, it has been shown that the presence of transition metals decreases the
typical temperature required for tube production (e.g., Jung et al., *Diamond and*
Related Materials 2001, **10**, 1235; Govindaraj et al., *Materials Research Bulletin*
1998, **33**, 663; Shyu and Hong, *Diamond and Related Materials* 2001, **10**, 1241).
35 Since their discovery, several different production methods have been introduced to

synthesize CNTs. These methods can be broadly divided into chemical and physical according to the method applied in releasing carbon atoms from carbon-containing precursor molecules. In the physical methods, e.g. arc-discharge (Iijima, *Nature* 1991, **354**, 56) and laser ablation (Guo et al., *Chem. Phys. Lett.* 1995, **243** 49), high-energy input is used to release the carbon atoms needed for tube synthesis. The chemical methods rely on carbon atomization via catalytic decomposition of carbon precursors on the surface of transition metal particles. According to the place where the growth of CNTs occur, chemical methods for their production can be divided into surface supported, so called CVD (chemical vapor deposition) [e.g. Dai et al., *Chem. Phys. Lett.* 1996, **260**, 471] and aerosol [e.g. Bladh, Falk, and Rohmund, *Appl. Phys. A*, 2000, **70** 317; Nasibulin et al., *Carbon*, 2003, **41**, 2711] syntheses. In CVD methods, the carbon precursor decomposition and CNT formation take place on the surface of catalyst particles that are supported on a substrate. In aerosol synthesis, the catalyst particles are in the gas-phase. The terms “gas-phase synthesis” and “floating catalyst method” have been also applied in the literature for this process. We use the term “aerosol synthesis” to specify processes taking place completely in the gas-phase.

The method described in this patent is a new approach to the production of single walled and multi-walled CNTs, functionalized CNTs and CNT composite materials and matrices thereof. This new method requires pre-made catalyst particles or a procedure to produce pre-made catalyst particles with a narrow distribution of properties, a carbon source, a reagent, when needed, an energy source, when needed and a flow control system. A principle advantage of the new method over existing methods is that it allows the separate control of the introduction of catalyst particles and the CNT synthesis. In other methods, catalyst particles are formed by gaseous chemical reactions leading to the formation of supersaturated vapor of the catalyst material (e.g. WO 00/26138) or physical nucleation directly from supersaturated gas (e.g. WO 03/056078) simultaneously with the CNT synthesis and thus cannot be separately controlled. This leads to the formation of CNTs with potentially large variation in important properties such as length, diameter and chirality. The diameter and chirality of the CNTs produced via catalysts are largely determined by the properties of the catalyst particles, in particular the catalyst size. Though patent US 2002/102193 A describes a means of separately producing catalyst particles and CNTs, it does not specify a means of

controlling the high non-uniformity of catalyst particles produced by the chemical nucleation method proposed and thus will tend to produce non-uniform CNTs. Our invention, on the other hand, provides a means of separately introducing catalyst particles with well controlled properties, either directly through a process which
5 inherently produces catalysts with narrow particle size distributions (e.g. the physical vapor nucleation processes described in this method), or by providing specific means of narrowing the size distribution from processes (such as the chemical nucleation method referred to in US 2002/102193 A) which inherently produce wide catalyst particle size distributions and thus non-uniform CNTs. As the
10 industrial and scientific utility of produced CNTs is a function of their individual and collective properties, there exists an urgent need for CNTs and a method for production of CNTs and CNT composite formulations with more uniform and controlled properties.

In our method we utilize pre-made particles for production of CNTs and
15 CNT composite formulations. Those pre-made particles can be prepared by conventional methods such as chemical vapor decomposition of catalyst precursor [e.g. Nasibulin et al., *J. Phys. Chem. B*, 2001, **105**, 11067.], by the physical vapor nucleation method, which implies an evaporation and subsequent vapor nucleation followed by growth of particles due to vapor condensation and cluster coagulation
20 (for instance, a resistively heated hot wire generator, an adiabatic expansion in a nozzle or an arc discharge method), by thermal decomposition of precursor solution droplets (e.g. by electrospray thermal decomposition) or by any available method which either inherently produces particles with a narrow distribution of properties or can be pre-classified prior to CNT synthesis to narrow the distribution. The pre-
25 made particles are then introduced into a CNT reactor where CNT synthesis takes place. Thus, the current invention separates the catalyst production from the CNT synthesis and allows the control of each step in the production process. In order to produce CNTs with further controlled properties, the pre-made particles, either produced as part of the process or introduced from existing sources, can be classified
30 according to size, mobility, morphology or other properties before being introduced into one or more CNT reactors. Furthermore, the current invention allows the continuous or batch production of composite CNT either coated or mixed with additive materials. Additionally, the current invention provides a means of producing pure, functionalized or composite CNT gas, liquid or solid dispersions,

solid structures, powders, pastes, colloidal suspensions and surface depositions and can be integrated directly into a means of fabricating structures from such materials. Additionally, when used in conjunction with the physical nucleation method, the current invention provides the additional advantage of allowing better control over
5 conditions in the CNT reactor conditions since physical nucleation introduces no additional chemical compounds into the environment which can interfere with CNT formation, growth, purification and/or functionalization.

2. SUMMARY OF THE INVENTION

The present invention relates to single walled and multi-walled carbon
10 nanotubes (CNTs), functionalized carbon nanotubes and carbon nanotube composites with controlled properties, to a method for aerosol synthesis of single walled and multi-walled carbon nanotubes, functionalized carbon nanotubes and carbon nanotube composites with controlled properties from pre-made catalyst particles and a carbon source in the presence of zero or more reagents and zero or
15 more additives, to functional, matrix and composite materials composed thereof and structures and devices fabricated from the same in one or more continuous or batch CNT reactors. This method comprises the steps of:

- (a) formation of catalyst particles (so-called pre-made particles), if needed;
- (b) size classification of the pre-made catalyst particles, if needed;
- 20 (c) introducing the pre-made catalyst particles into the CNT reactor;
- (d) introducing one or more carbon sources into the CNT reactor;
- (e) catalytic decomposition of one or more carbon sources;
- (f) formation of CNTs;
- (g) introducing zero or more reagents, which can be done together with
25 carbon sources or separately before, during or after the CNT formation, to promote CNT formation, to purify CNTs, to dope CNTs, and/or to functionalize the produced CNTs when desired;
- (h) introducing zero or more additives to the CNT aerosol to produce a CNT composite material when desired;
- 30 (i) collection of produced CNTs and/or CNT formulations in a solid, liquid or gas dispersion, a solid structure, a powder, a paste, a colloidal suspension and/or as a surface deposition when desired;
- (j) deposition of gas dispersions of produced CNTs and/or composite CNT formulations onto surfaces and/or into matrix and/or layered structures
35 and/or devices when desired.

The present invention includes one or more CNT reactors, which can allow continuous or batch production of CNTs, functionalized CNTs, doped CNTs and

composites thereof. The present invention allows all or part of the processes of synthesis of CNTs, their purification, doping, functionalization, coating, mixing and deposition to be combined in one continuous procedure and in which the catalyst synthesis, the CNT synthesis, and their functionalization, doping, coating, mixing
5 and deposition can be separately controlled. The present invention further provides a composition of matter comprising single walled and multi-walled CNTs and structures and devices fabricated from the same.

3. BRIEF DESCRIPTION OF DRAWINGS

10 **Figure 1** shows a block diagram of an arrangement for the method for CNT production.

Figure 2 shows thermodynamic calculations of free Gibbs' energy of decompositions of example carbon sources at different temperatures.

Figure 3 shows a preferred embodiment of the invention for CNT production
15 where the pre-made catalyst particles were formed by a physical vapor nucleation method from a hot wire generator (a) separated in space from the CNT reactor and (b) smoothly integrated with the CNT reactor.

Figure 4 shows CFD calculations of temperature contours in the vicinity of a resistively heated wire (Inflow velocity $U=1$ m/s, $T_{\text{gas}}=273$ K, $T_{\text{wire}}=1273$ K.
20 Gravity points to the left).

Figure 5 shows CFD calculations of (a) the temperature profile and (b) velocity vectors in a preferred embodiment of the invention. (Maximum wall $T_{\text{wall}}=1273$ K, inner flow rate = 0.4 LPM, outer flow rate = 0.8 LPM. Gravity points to the left).

25 **Figure 6(a)** shows an alternate embodiment of the invention for production of single walled and multi-walled CNTs, where the pre-made catalyst particles are formed by decomposing one or more catalyst particle precursors.

Figure 6(b) shows an alternate embodiment of the invention for production of single walled and multi-walled CNTs, where the pre-made catalyst particles are
30 formed by a physical vapor nucleation method (for instance, by an arc discharge) or by an electrospray thermal decomposition method.

Figure 6(c) shows an alternate embodiment of the invention for batch production of pre-made catalyst particles in combination with continuous production of single walled and multi-walled CNTs and CNT composites.

35 **Figure 6(d)** shows an alternate embodiment of the invention for batch production of pre-made catalyst particles in combination batch production of single walled and multi-walled CNTs and CNT composites.

Figure 6(e) shows an alternate embodiment of the invention for a single batch CNT reactor for production of pre-made catalyst particles and single walled and multi-walled CNTs and CNT composites.

5 **Figure 6(f)** shows an alternate embodiment of the invention for continuous production of CNTs wherein sheath gas is used to insure catalyst particles and CNTs are not deposited on CNT reactor walls, thus avoiding surface growth of CNTs and CNT composites.

10 **Figure 6(g)** shows an alternate embodiment of the invention for continuous production of CNTs wherein a controlled temperature gradient in the CNT reactor is used to separate catalyst particle synthesis from CNT synthesis.

Figure 6(h) shows an alternate embodiment of the invention for production of composite CNTs wherein an additional flow of coating material or particles is introduced into the CNT aerosol flow and mixed to create a composite formulation.

15 **Figure 6(i)** shows a CFD calculation of an alternate embodiment of the invention for production of CNTs and/or CNT composite formulations wherein controlled sampling of the product aerosol is used to isolate a portion of the aerosol flow that has experienced essentially uniform conditions as it has passed through the reactor(s) and/or pre-reactor(s).

20 **Figure 7** shows TEM images of single walled CNTs synthesized at 1200 °C from carbon monoxide as a carbon source using iron as a catalyst material.

Figure 8 shows TEM and SEM images of multi-walled CNTs scratched from the CNT reactor walls.

25 **Figure 9(a)** shows thermodynamic calculations of the temperature dependence of mol fraction of the product after mixing 1 mol of CO and 1 mol of H₂.

Figure 9(b) shows thermodynamic data for CO disproportionation: dependencies of free energy change, ΔG , and CO mol fraction in gaseous phase on temperature. Kinetic data: CO concentration after disproportionation on surface of nanometer iron particles.

30 **Figure 9(c)** shows thermodynamic calculations of the temperature dependence of the mol fraction of hydrogen atoms.

Figure 9(d) shows thermodynamic temperature dependencies of the free energy change, ΔG , of reaction, leading to the liberation of carbon.

35 **Figure 10(a)** shows number length distributions of CNTs produced at various conditions.

Figure 10(b) shows number diameter distributions of CNTs produced at various conditions.

Figure 10(c) shows the correlation between diameters of catalyst particles and produced CNTs.

Figure 10(d) shows the dependence of the length of produced CNTs on the CO flow rate.

Figure 11 shows TEM images of single walled CNTs synthesized inside a stainless steel tube at a set furnace temperature of 900 °C ($t_{max} = 1070$ °C).

5 **Figure 12** shows TEM images of single walled CNTs synthesized at 1200 °C from carbon monoxide and ethanol/thiophene mixture as carbon sources and reagents using nickel as a catalyst material.

Figure 13 shows TEM images of single walled CNTs synthesized at 1200 °C from ethanol as a carbon source and a reagent using iron as a catalyst material.

10 **Figure 14** shows TEM images of single walled CNTs synthesized at 1200 °C from ethanol/thiophene mixture as carbon sources and reagents using iron as a catalyst material.

15 **Figure 15** shows TEM images of multi-walled CNTs synthesized at 1200 °C from ethanol/thiophene mixture as carbon sources and reagents using iron as a catalyst material.

Figure 16 shows TEM images of fullerene functionalized CNTs synthesized at 900°C from CO as carbon source and pure hydrogen as a reagent flowing through the hot wire generator using iron as a catalyst material and using a stainless steel reactor tube.

20 **Figure 17** shows TEM images of fullerene functionalized CNTs synthesized at 900°C from CO as carbon source and water vapor as a reagent and using iron as a catalyst material and using a stainless steel reactor tube.

25 **4. DETAILED DESCRIPTION OF THE INVENTION**

The present invention relates to single walled and multi-walled carbon nanotubes (CNTs), functionalized carbon nanotubes and carbon nanotube composites with controlled properties, to a method for aerosol synthesis of single walled and multi-walled carbon nanotubes, functionalized carbon nanotubes and carbon nanotube composites with controlled properties from pre-made catalyst particles and a carbon source in the presence of zero or more reagents and zero or more additives, to functional, matrix and composite materials composed thereof and structures and devices fabricated from the same in one or more continuous or batch CNT reactors. Once the CNTs are formed, they can be purified, further functionalized and/or doped and/or further coated with additional materials by, for instance, condensation of supersaturated gas on the CNT surface or by being mixed with an additional aerosol source, thus creating composite CNTs.

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Figure 1 shows a block diagram of an arrangement of the invention for single walled and multi-walled CNT production. The method can be a continuous flow, batch or a combination of batch and continuous sub-processes. The first step of the method is to obtain aerosolized pre-made catalyst particles. These particles can be produced as part of the process or can come from an existing source. Those particles can be classified according to important properties (for instance, size, mass, shape, crystallinity, charge or mobility) or, when the distribution of properties is sufficiently narrow, can be directly introduced into the CNT reactor. In the CNT reactor, the pre-made catalyst particles are mixed and heated together with one or more carbon sources and with zero or more reagents. Then, the carbon source catalytically decomposes. Reagents can be added into the CNT reactor for chemical reaction with catalyst particles and/or carbon source and/or with CNTs. Thus, reagents can be added together with one or more carbon sources, after the decomposition of the carbon source, and/or after the CNT formation. During or after the formation of CNTs, the entire product or some sampled portion of the product can be selected for further processing steps such as functionalization, purification, doping, coating and mixing. All or a sampled part of the resulting raw CNT product can then be collected directly, or incorporated into a functional product material which can further be incorporated in devices.

20 ***Carbon sources***

According to the present invention, as a carbon source, various carbon containing precursors can be used. Carbon sources include, but are not limited to, gaseous carbon compounds such as methane, ethane, propane, ethylene, acetylene as well as liquid volatile carbon sources as benzene, toluene, xylenes, trimethylbenzenes, methanol, ethanol, and/or octanol. Alternatively and preferably, carbon monoxide gas alone or in the presence of hydrogen can be used as a carbon source. Other carbon sources are possible and these examples are not intended to limit the scope of the invention in any way.

Thermodynamic calculations for decompositions of various carbon sources are presented in Figure 2. This figure is provided to illustrate the variety of possible desirable carbon sources and does not, in anyway, limit the sources for which the present method can be applied. Note that increasing the temperature in the system generally makes the potential carbon sources less stable, except for the reactions connected with carbon monoxide: CO disproportionation and the reaction between CO and H₂. Those reactions will be discussed further as related to the description of the preferred embodiment and in Example 1.

For other carbon sources, increasing the number of atoms in saturated hydrocarbons (CH₄, C₂H₆, C₃H₈) generally leads to a decrease in the stability of the

substances. The stability behavior of systems with saturated carbon bonds from C_2H_2 via C_2H_4 to C_2H_6 is more complex, because of its complex temperature dependency. Some liquid carbon sources are also included in this figure. One can see a trend for aromatic compounds (benzene C_6H_6 , toluene $C_6H_5-CH_3$, o-xylene $C_6H_4-(CH_3)_2$, 1,2,4-trimethylbenzene $C_6H_3-(CH_3)_3$). The most stable is the benzene molecule. Increasing the amount of methyl groups in the compound makes compounds less stable. To control the properties of produced CNTs, such as chirality, fullerene molecules can be also used as a carbon source. Nevertheless, all of the presented compounds and many other carbon containing molecules can be used as a carbon source in the present invention. It is worth noting that the decomposition of carbon sources can occur even without the presence of catalyst particles, but because the decomposition is a kinetically limited process, a reasonable decomposition rate at moderate temperatures and relatively low residence times can be obtained in the presence of catalyst particles.

Additionally, other methods can be used to activate carbon precursors at desired locations in the reactors by using, for instance, heated filaments.

Catalyst Particles

As a catalyst material, various transition metals, which catalyze the process of carbon source decomposition/disproportionation known in the art can be used. A preferred catalyst particle consist of transition metals and combinations thereof, but other materials are possible. Generally preferred for CNT production are catalyst based on iron, cobalt, nickel, chromium, molybdenum, palladium. Other metal and non-metal materials are possible according the invention and the preceding examples are not intended to limit the scope of the invention in any way.

The catalyst particles to be introduced into the CNT reactor can be produced by various methods known in the art such as chemical vapor decomposition of catalyst precursor, physical vapor nucleation, or of droplets made by electrospray, ultrasonic atomization, air atomization and the like or thermal drying and decomposition. Other methods for producing catalyst particles are possible according to the invention and the preceding list is in no way intended to limit the processes applicable. Additionally, pre-made catalyst particles can be synthesized in advance and then introduced into the CNT reactor, though, generally, particles of the size range needed for CNT production are difficult to handle and store and thus it is preferable to produce them in the vicinity of the CNT reactor as an integrated step in the CNT and composite CNT production process.

For the chemical method of pre-made catalyst particle production, metalorganic, organometallic or inorganic compounds such as metallocene,

carbonyl, and chelate compounds known in the art can be used as catalyst precursors. In general, however, due to the relatively slow decomposition reaction for these precursors, relatively wide particle size distributions are achieved with these methods, thus, to achieve the desired control of catalyst particle sizes, these methods should be used in conjunction with a pre-classifier. Other compounds are possible according to the invention and these examples are in no way intended to limit the compounds available according to the invention.

For the physical method of pre-made catalyst particle production, pure metals or their alloys can be evaporated by using various energy sources such as resistive, conductive or radiative heating or chemical reaction (wherein the concentration of produced catalyst vapor is below the level needed for nucleation at the location of release) and subsequently nucleated, condensed and coagulated from supersaturated vapor. Means of creating supersaturated vapor leading to the formation of catalyst particles in the physical method include gas cooling by convective, conductive and/or radiative heat transfer around, for instance, a resistively heated wire and/or adiabatic expansion in, for instance, a nozzle. The hot wire method developed here, however, is preferable in that it inherently produces catalyst particles with a narrow size distribution and thus does not require a pre-classification step to produce CNTs with a narrow distribution of properties.

For the thermal decomposition method of pre-made catalyst particle production, inorganic salts can be used such as nitrates, carbonates, chlorides, fluorides of various metals. Other materials are possible according to the present invention and these examples are not intended to limit the scope of the invention in any way

In order stabilize the CNT production and to enhance the decomposition of carbon precursor at the catalyst particle surface, the reactor walls preferably, but not necessary, should be saturated by the catalyst material. The wall saturation can be done by any available methods. As an example, a chemical vapor deposition of the material on walls using a catalyst material compound as a precursor or by evaporation of catalyst material and allowing its condensation on the reactor walls can be done. Another possibility to have saturated conditions is to use a reactor tube made of the catalyst containing material as is demonstrated in Example 3.

For the production of CNTs with further controlled properties, the pre-made particles can be classified according to, for instance, mobility or size and by, for

instance, differential mobility analyzers (DMA) or mass spectrometers. Other methods and criteria for classification are possible according to the present invention and the preceding examples are not intended to limit the scope of the invention in any way. Additionally, flow control, with, for instance, sheath gas and isokinetic
5 sampling can be used to, for instance, provide uniform conditions for particle and CNT formation and growth and/or to classify product according to reactor conditions so as to achieve for uniform CNT and composite CNT properties.

Promotion, Purification, Functionalization and Doping of CNTs

10 Reagents are needed for participation in the chemical reaction with catalyst particle precursor and/or with catalyst particles and/or with carbon source and/or with amorphous carbon and/or with CNTs. The purpose of the reagent is to be a promoter for the CNT formation and/or to increase (or decrease) the rate of carbon source decomposition and/or to react with amorphous carbon during or after the
15 production of CNTs for purification and/or to react with CNTs for functionalization and/or doping of CNTs. The reagents can also behave as a carbon source according to the present invention.

As a promoter for CNT formation, preferably sulfur, phosphorus or nitrogen elements or their compounds such as thiophene, PH_3 , NH_3 can be used. Additional
20 promoters include H_2O , CO_2 and NO . Other promoter compounds known in the art are possible according to the present invention and these examples are not intended to limit the scope of the invention in any way.

Purification processes are generally needed to remove undesirable amorphous carbon coatings and/or catalyst particles encapsulated in CNTs. Usually this
25 procedure takes significant time and energy, often more than required for the CNT production itself. In the present invention it is possible to have one or more separated heated CNT reactors/reactor sections, where one CNT reactor or section of the CNT reactor is used to produce CNTs and the other(s) are used for, for instance, purification, functionalization and/or doping. It is also possible to combine the
30 growth and functionalization steps as shown in examples 7 and 8. Amorphous carbon, deposited on the surface of CNTs, can be removed in one or more subsequent CNT reactors/reactor sections by, for instance, heat treatment and/or addition of special compounds which, for instance, form reactive radicals (for instance, OH), which react with undesirable products rather than with CNTs. One or
35 more subsequent CNT reactors/sections can be used for, for instance, the removal of catalyst particles from the CNTs by creating the conditions where the catalyst particles evaporate as was shown in [Nasibulin et al., *Carbon* 2003, 412, 2711 and

FI-20035120]. Other processing steps are possible according to the present invention.

As a reagent for the reaction with a carbon source to alter its decomposition rate, hydrogen can be used. As an example, carbon monoxide reacts with hydrogen, namely, with hydrogen atoms, which are formed at high temperatures due to decomposition of hydrogen molecules.

As a chemical for amorphous carbon removal, any compounds or their derivatives or their decomposition products formed *in situ* in the CNT reactor, which preferably react with amorphous carbon rather than with graphitized carbon, can be used. As an example of such reagents known in the art, alcohols, ketones, organic and inorganic acids can be used. Additionally, oxidizing agents such as H_2O , CO_2 or NO can be used. Other reagents are possible according to the present invention and these examples are not intended to limit the scope of the invention in any way.

Another role of the reagent is to functionalize the CNTs. Chemical groups attached to CNTs alter the properties of the produced CNTs. Functionalization and doping of CNTs can radically change such properties as solubility and electronic structure (varying from wide band gap via zero-gap semiconductors to CNTs with metallic properties). As an example, the doping of CNTs by lithium, sodium, or potassium elements leads to the change of the conductivity of CNTs, namely, to obtain CNTs possessing superconductive properties. Functionalization of CNTs with fullerenes produces semi-conducting CNTs and allows further functionalization of the CNTs via the attached fullerenes by methods known in the art. In the current invention, the in-situ functionalization and/or doping can be achieved via the introduction of appropriate reagent before, during or after CNT formation.

Moreover, the reagent, which can be used for promotion, purification, functionalization, and/or doping of CNTs can be a carbon source as well. Also a carbon source, which can be used for the CNT production, can also be a reagent.

CNT Composites by Coating and Mixing

One or more additives can be used for coating and/or mixing with the produced CNTs to create composite CNT formulations. The purpose of the additives are to, for instance, increase the catalytic efficiency of particles deposited in a matrix or to control matrix properties of such as hardness, stiffness and thermal and electrical conductivity or expansion coefficient. As a coating or particle additive for CNT composite materials, preferably one or more metal containing or organic materials such as polymers or ceramics can be used. Other additive compounds are possible according to the present invention and these examples are not intended to limit the scope of the invention in any way. These can be deposited as a surface coating on the CNTs through, for instance, condensation of supersaturated vapor,

chemical reaction with previously deposited layers, doping agents or functional groups or by other means known in the art or, in the case that the additive is a particle, mixed and agglomerated in the gas phase. Additionally, gas and particle deposition on CNTs can be combined.

5 **Classification**

In order to produce CNTs with further controlled properties, pre-made particles, either produced as part of the process or introduced from existing sources, can be classified according to size, mobility, morphology or other properties before being introduced into the CNT reactor(s), where the CNT formation occurs. For
10 instance, a high resolution differential mobility analyzer (HR-DMA) [Nasibulin et al., *J. Nanoparticle Res.* 2002, 4, 449], which allows a very high resolution particle size selection with a standard deviation of $\sigma \leq 1.025$ at 1 nm particle size, can be used as a classifier. Other examples include, but are not limited to, mass spectroscopy, sedimentation, diffusion, centrifugation, solvation, and chemical
15 reaction according to the invention. Additionally, controlling the flow field and temperature distribution in the reactor can be used as a means to control and/or classify catalyst particle properties.

Energy Sources

Various energy sources can be used, when desired, to promote or impede, for
20 instance, chemical reactions and CNT synthesis according to the invention. Examples include, but are not limited to, resistively, conductively, radiatively or nuclear or chemical reactively heated CNT reactors and/or pre-reactors.

Controlled Sampling and Deposition of Aerosol Product

Various means can be used, when desired, to control or selectively sample
25 the CNT and composite CNT before and/or after functionalization, purification, coating, mixing and/or doping. Such control devices reduce the variation of product properties by selecting only those products that have been exposed to similar environmental conditions. Various means of controlled sampling of the aerosol product are possible according to the invention, including, but not limited to,
30 selective sampling from regions of the reactor with uniform conditions and aerosol focusing through particle lenses, acoustic focusing devices, and electrical focusing fields. Similarly, these techniques can be combined by those experienced in the art to further enhance their control effects.

Controlled deposition of synthesized materials can be achieved by various means including, but not limited to inertial impaction, thermophoresis and/or migration in an electrical field to form desired geometries (e.g. lines, dots or films) with desired properties such as electrical or thermal conductivity, opacity or mechanical strength, hardness or ductility.

5. Detailed description of a preferred embodiment of the invention

Figure 3(a) shows the preferred embodiment of the invention for the continuous production of single walled or multi-walled CNTs where the pre-made catalyst particles are formed by the physical vapor nucleation method from a hot wire generator (HWG) (3) separated in space from the CNT reactor. In said embodiment, a carbon source is supplied either from a carrier gas reservoir (1) (e.g. carbon monoxide, methane, ethane, etc.) or by a carrier gas passing through a saturator (6). The saturator can also be used to introduce reagents for e.g. CNT purification and/or functionalization. If the carbon source is a solid substance, it can be heated to increase the equilibrium vapor pressure. For liquid substances, the saturator can be, for instance, a bubbler. Room temperature is a suitable temperature to provide a suitable vapor pressure for some liquid carbon precursors (for instance, for methanol, ethanol, octanol, benzene, toluene, etc.). Nevertheless, the vapor pressure of the liquid substance can be adjusted by heating or cooling the bubbler or by dilution.

Another carrier gas (pure nitrogen or nitrogen/hydrogen mixture, 93% / 7%) is supplied from a carrier gas reservoir (2) to the HWG (3), which is operated with the help of an electric power supply (4). As the carrier gas passes over the heated wire, it is saturated by the wire material vapor. After passing the hot region of the HWG, the vapor becomes supersaturated, which leads to the formation of pre-made particle due to the vapor nucleation and subsequent vapor condensation and cluster coagulation. Inside the CNT reactor (5) or before, when needed, the two separate flows containing the pre-made catalyst particles and the carbon source are mixed and subsequently heated to the CNT reactor temperature. The carbon source can be introduced through the HWG if it does not react with the wire. Other configurations are possible according to the invention, so long as the catalyst particles are formed before CNT synthesis begins.

It is known that nanoparticles possess very high diffusivity and high pinning energy with surfaces. In order to avoid diffusion losses of the catalyst particles and to use them more efficiently, the distance between the HWG and the location where the formation of CNT occurs, can be adjusted. Figure 3(b) shows the equivalent embodiment when the pre-made catalyst particles are formed by a physical vapor

nucleation method from a hot wire generator smoothly integrated with the CNT reactor. Here, the HWG is located inside the first section of the CNT reactor. In this preferred embodiment, the end of the HWG tube was placed at the location where the CNT reactor wall temperature of about 400 °C. This temperature was found to be optimal, since reduced particle growth due to the catalyst particle agglomeration and coagulation, minimized particle diffusion losses on the walls and provided a reasonable iron vapor nucleation rate.

The metal particle size is of great importance in the formation of CNTs since CNT diameter has been shown to correlate with catalyst particle size. The nucleation rate and final particle size depend on the temperature gradient over the metal wire and on the concentration of the metal vapor. The concentration of the vapor and the temperature gradient on the other hand depend on the gas flow rate over the metal wire and the wire temperature. Since large temperature gradients (≈ 500000 K/s) can be achieved, the HWG can be applied to the production of very small primary particles. The temperature change over the heated metal wire was calculated with a Computational Fluid Dynamics (CFD) model. In the calculation, an incoming gas velocity (U) of 1 m/s and temperature of 273 K were used. As can be seen in Figure 4, the temperature gradient near the wire surface is extremely large meaning that the metal vapor rapidly cools down (approximately 500 °C in 1 mm distance). Correspondingly, the vapor reaches supersaturation very quickly, which in turn results in homogeneous nucleation of large number of small metal clusters. Calculations show that the temperature drops such that homogenous nucleation of catalyst particles should be complete with a few millimeters of the hotwire. Furthermore, it has been found that the method produces exceptionally narrow particle size distributions and so can be used in the current invention without the necessity of a particle classification step as would be needed in, for instance, typical chemical nucleation methods.

CFD calculations were carried out to define the temperature and velocity profiles and mixing conditions in the CNT reactor (namely, in the preferred embodiment shown in Figure 3(b)) under laminar conditions including the effects of buoyancy. Results of the CFD calculations are shown in Figure 5(a) and Figure 5(b) and exhibit how the current invention can be constructed to define the residence time and temperature history of carrier gases and reagents, catalyst particles and carbon nanotubes in the CNT reactor so as to control catalyst particle and nanotube growth.

6. Description of sample alternate embodiments.

Figure 6(a) shows another embodiment used, according to the present invention, for the production of the single walled and multi-walled CNTs. In this

figure, the system for production of pre-made catalyst particles consist of a carrier gas cylinder (2), saturator (8) and (6), a pre-reactor (7) and a particle classifier (9). It should be noted that the carrier gas can be a carbon source as well. The saturator (8) can be used for the carrier gas saturation by a carbon source. The saturator (6) can be used for the carrier gas saturation by a catalyst precursor. Saturators (6) and (8) can also be used to introduce reagents into the system for, for instance, CNT purification or functionalization. If the catalyst precursor and carbon source are solid substances, they can be heated to increase their equilibrium vapor pressures. For liquid substances, the saturator can be, for instance, a bubbler. Room temperature is a suitable temperature for a necessary vapor pressure of some liquid catalyst precursors (for instance, for iron pentacarbonyl) and carbon sources (for instance, benzene and toluene). Nevertheless, the vapor pressure of the liquid substance can be adjusted by heating or cooling the bubbler. Another possibility to decrease the vapor pressure of the liquid after the bubbler is to dilute the liquid with a suitable solvent or to dilute the vapor with an inert gas. For instance, a mixture of benzene and cobalt carbonyl can be used to decrease the vapor pressure of $\text{Co}(\text{CO})_4$. Moreover, one or more furnaces or furnace sections can be used. Zero or more furnaces/furnace sections can be used for catalyst production and one or more furnaces/furnace sections can be used for CNT formation. Additional furnaces/furnace sections can be used for purification and/or functionalization and/or doping of CNTs. Zero or more reagents can be added in the system before, during and/or after CNT formation.

The pre-reactor (7) and/or CNT reactor (5) can be, but are not necessarily, resistively heated. Other energy sources can be applied to energize and decompose the precursor. For instance, it can be radio-frequency, microwave, acoustic, laser induction heating or some other energy source such as chemical reaction.

The formed pre-made catalyst particles can be classified in size in a particle classifier (9). For this purpose, a differential mobility analyzer can be used. Other criteria and methods can, according to the invention, be used for classification. Subsequently, the pre-made particles are introduced into the CNT reactor.

A sample alternate embodiment of the invention for continuous single walled and multi-walled CNT production where the pre-made catalyst particles are made by the physical vapor nucleation method (for instance, adiabatic expansion in a nozzle or an arc discharge) or by thermal decomposition of precursor solution droplets is shown in Figure 6(b). All the elements remain the same as in the previous sample alternate embodiment except that, instead of the saturator (8) and the pre-reactor (7) (in Figure 6(a)), another system for the production of the pre-made particles (10). Box (10) depicts, for example, adiabatic expansion in a nozzle, an arc discharge or electrospray system for the formation of metal containing particles. Other methods

are applicable according to the invention and these examples are not intended to limit the scope of the invention in any way. Box (10) can also represent a means of aerosolizing pre-existing catalyst particles. The aerosol pre-made particles can be classified in a classifier (9) or introduced directly to the CNT reactor (5). Methods involving chemical nucleation will, in general, require pre-classification to achieve the desired uniformity in particle properties for well controlled CNT production.

A sample alternate embodiment of the invention for batch production of pre-made particles and continuous production of single walled and multi-walled CNT production is shown in Figure 6(c). As in the continuous processes, the pre-made particles can be prepared by any of the described methods such as physical nucleation, chemical vapor decomposition, or electrospray thermal decomposition in one or more batch CNT reactors (11) by introducing one or more carriers, catalyst precursors, carbon sources and/or reagents through one or more inlet(s)/outlet(s) (12) and subsequently evacuated after the batch process is completed through inlet(s)/outlet(s) (12). Alternately, pre-made catalyst particles can be directly introduced into the CNT reactor(s) or first classified in the classifier (9).

A sample alternate embodiment of the invention for batch production of pre-made particles and batch production of single walled and multi-walled CNT production is shown in Figure 6(d). As in the continuous processes, the pre-made particles can be prepared by any method such as physical nucleation, chemical vapor decomposition, or electrospray thermal decomposition in one or more batch CNT reactors (11) by introducing one or more carriers, catalyst precursors, carbon sources and/or reagents through one or more inlet(s)/outlet(s) (12) and subsequently evacuated after the batch process is completed through inlet(s)/outlet(s) (12). Alternately, pre-made catalyst particles can be directly introduced into the CNT reactor(s). Once these particles are produced, they can be introduced into the CNT reactor (13) through one or more inlet(s)/outlet(s) (14) where the time, gas composition and temperature history can be adjusted for CNT growth. Subsequently, the CNT reactor can be evacuated through inlet(s)/outlet(s) (14) and the CNTs collected.

Figure 6(e) depicts a sample embodiment wherein only one batch CNT reactor is used for both production of pre-made catalyst particles and for CNT synthesis. As in the previous embodiments, the pre-made particles can be prepared by method such as physical nucleation, chemical vapor decomposition, or electrospray thermal decomposition in a batch CNT reactor (13) by introducing precursors, reagents and/or carrier gases through one or more inlet(s)/outlet(s) (14). Alternately, pre-made catalyst particles can be directly introduced into the CNT reactor(s). Once the batch process is complete, appropriate carriers, catalyst precursors, carbon sources and/or reagents are introduced into the CNT reactor (13)

through one or more inlet(s)/outlet(s) (14) where the time, gas composition and temperature history can be adjusted for CNT growth. Subsequently, the CNT reactor can be evacuated through inlet(s)/outlet(s) (14) and the CNTs collected.

Figure 6(f) depicts a sample embodiment wherein sheath gas is used to control the catalyst particle and CNT deposition and heating in the CNT reactor tube in the case of a continuous flow system. Here a furnace (17) heats one or more carriers, catalyst particles, catalyst precursors, carbon sources and/or reagents introduced through inlet (18). Additional sheath gases are fed to the CNT reactor through one or more porous tubes (21), thus insuring the CNT reactor surfaces are free of catalyst particles and CNTs. Said sheath flow(s) can consist of one or more carriers, catalyst precursors, carbon sources and/or reagents according to the invention. The resulting aerosol then exits the CNT reactor through outlet (19). Other methods of flow control to minimize catalyst particle and CNT deposition are possible according to the invention.

Figure 6(g) depicts a sample embodiment wherein a single furnace with a gradually increasing wall temperature is used to separate the catalyst particle production from the CNT formation. In this embodiment, a continuous flow CNT reactor is divided into multiple temperature heating blocks (22) and (23). All required carrier gases, catalyst precursors, carbon sources and/or reagents are introduced through inlet (18). The temperature of heating block (18) is set high enough such that the catalyst particle precursor decomposes to produce catalyst particles by a chemical nucleation method but below that needed to initiate CNT synthesis. The temperature of heating block (22) is set above that needed to initiate CNT synthesis. Each block of the CNT reactor can then be controlled independently thus creating two distinct CNT reactor sections smoothly integrated with one another. Other methods of separating the catalyst particle synthesis and CNT synthesis in a continuous or batch production process are possible according to the invention.

Figure 6(h) depicts a sample embodiment of the invention for continuous production of CNT composites wherein an additional flow of additive coating material or aerosolized particles (24) is introduced into the CNT aerosol flow (25) to create a composite material. Examples of possible additives include, but are not limited to, polymers, metals, solvents and ceramics and aerosols thereof. The resulting composite aerosol (26) can then be directly collected, deposited in a matrix or deposited on a surface by electrical, thermophoretic, inertial, diffusional, turbophoretic, gravitational or other forces known to the art to form thick or thin films, lines, structures and/or layered materials. Further control can be achieved by, for instance, jet focusing of the resulting CNT aerosol stream.

Figure 6(i) shows a CFD calculation of an alternate embodiment of the invention for production of CNTs and/or CNT composite materials wherein controlled sampling of the product aerosol is used to isolate a portion of the aerosol flow that has experienced essentially uniform conditions near the reactor centerline throughout the reactor(s) and/or pre-reactor(s). Other means of controlled sampling of the aerosol product are possible according to the invention, including, but not limited to, aerosol focusing through particle lenses, acoustic focusing devices, and electrical focusing fields.

7. EXAMPLES

In order to facilitate a more complete understanding of the invention, examples are provided below. These examples are for purposes of illustration only and are not intended to limit the scope of the invention in any way.

In all the following examples, the morphology and the size of the product are investigated with a field emission transmission electron microscope (TEM, Philips CM200 FEG) and a field emission scanning electron microscope (Leo Gemini DSM982). Electron diffraction (ED) patterns of the products were used for determination of the crystalline phase of metal particles.

Where various embodiments of the present invention are described in detail, it is apparent that modifications and adaptations of those embodiments will occur to those skilled in the art. However, it is to be expressly understood that such modifications and adaptations are within the spirit and scope of the present invention.

Example 1. Single walled CNT synthesis from carbon monoxide as carbon source using iron as catalyst material and using a ceramic reactor tube

Carbon source: CO.

Catalyst particle source: hot wire generator.

Catalyst material: iron wire of 0.25 mm in diameter.

Operating furnace temperature: 1200 °C.

Operating flow rates: CO outer flow of 400 cm³/min and hydrogen/nitrogen (7/93) inner flow of 400 cm³/min.

This example, illustrating the synthesis of single walled CNTs, was carried out in the embodiment of the invention shown in Figure 3(b). Carbon monoxide was supplied from a gas cylinder (1) and the experimental setup did not contain a saturator (6). The embodiment consisted of a HWG smoothly integrated with a heated vertical tubular CNT reactor. A ceramic tube, with an internal diameter of 22 mm inserted inside the 90-cm length furnace (Entech, Sweden) was used as a CNT

reactor. Inside the CNT reactor another ceramic tube with external and internal diameters of 13 and 9 mm, respectively and with a length of 25 cm was inserted. The HWG, which was a resistively heated thin iron wire, was located inside the internal tube. The location of the internal tube could be adjusted. The end of the
5 HWG tube was placed at the location with the CNT reactor wall temperature of about 400 °C. This temperature was found to be optimal, since reduced particle growth due to the agglomeration and coagulation, minimized particle diffusion losses on the walls and provided a reasonable iron vapor nucleation rate.

In order to suppress the iron particle evaporation inside the reactor, the walls
10 of the reactor tube were saturated with iron by running HWG in nitrogen/hydrogen atmosphere without carbon monoxide. Also it is possible to saturate the reactor walls by blowing iron containing compound vapor through the heated up to about 1000 °C reactor. For this purpose, the vapor of ferrocene or iron pentacarbonyl can be used.

The metal particles produced by the HWG were carried into the CNT reactor
15 with nitrogen/hydrogen (with mol component ratio of 93.0/7.0) from gas cylinder (2) shown in Figure 3(b). In the CNT reactor, the flow of the metal particles from the HWG was mixed with the outer CO flow. Inside the CNT reactor CO disproportionation or hydrogenation took place on the surface of the formed metal particles. Downstream of the CNT reactor a porous tube dilutor (12 L/min) was
20 used to prevent the product deposition on the walls. The aerosol product was collected by an electrostatic precipitator (Combination electrostatic precipitator, InTox Products) on a carbon coated copper grid (SPI Lacey Carbon Grid). Figure 7 demonstrates the product formed at the given operating conditions. CNTs are single
25 walled. The number diameter and length distributions obtained on the basis of high-resolution TEM images are presented in Example 2. An important characteristic of this process is the efficiency of the catalyst material usage. Almost all catalyst particles initiated the growth of CNTs.

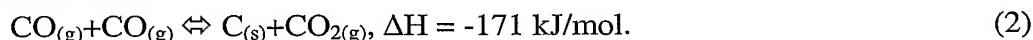
During the experiments also multiwalled CNTs (MWCNTs) were produced on the wall of the CNT reactor. Scratching the product from the walls upstream of
30 the CNT reactor at about 700 °C (after a 10 hour experiment) showed the presence of well crystalline MWCNTs among the product (Figure 8). The MWCNTs are shown to be a few microns long. TEM observations showed that the product consisted of different types of CNTs: bamboo-shaped tubes, MWCNTs with either a small (about 5) or a large (up to 50) number of walls. Scratching the product from
35 the walls after one week of operation showed very thick carbon tubes of about 200 nm in diameter. Thus, it is demonstrated that CNTs produced in the aerosol phase are markedly different from surface supported (i.e. CVD) produced CNTs.

Thermodynamic calculations

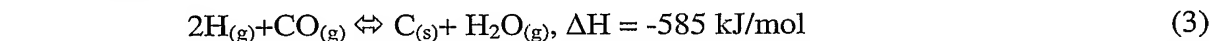
It is well known that at the studied furnace temperatures, two competing reactions, CO disproportionation and hydrogenation, leading to the formation of CNTs can occur. Since there were no CNT produced in the absence of hydrogen, we assume that the hydrogenation reaction of carbon monoxide



plays a very important role. The justification of the occurrence of this reaction can be seen from the thermodynamic calculations presented in Figure 9(a). It is worth noting that the concentration of the released carbon after reaction (1) is proportional to the concentration of water. Thus, reaction (1) can occur at temperatures lower than 900 °C, while at the temperatures higher than that, the reaction is prohibited. It is necessary to note that this behavior is similar to the reaction of CO disproportionation



In [Nasibulin et al., *Carbon*, 2003, **41**, 2711], aspects of the occurrence of this reaction was studied. As one can see from and Figure 9(b), reaction (2) is also inhibited at temperatures higher than about 900 °C and kinetic investigations showed an appreciable reaction rates in the temperature interval from 470 to 800 °C with a maximum rate at the temperature of 670 °C. It can be concluded that both reactions (1) and (2) occur in the same temperature range. The hypothesis about the leading role of reaction (1) in the formation of CNTs is supported by the fact that CNTs were produced only in the presence of hydrogen. The importance of hydrogen can be confirmed by the calculations presented in Figure 9(c) due to hydrogen molecule decomposition at the glowing wire temperatures. In the figure one can see a temperature dependence of equilibrium mol fraction of hydrogen atoms. The amount of hydrogen atoms is significant at the temperatures of the glowing iron wire (approx. 1500 °C). It is known that the formed hydrogen atoms are more reactive than the H₂ molecules. Moreover, a reaction between hydrogen atoms and carbon monoxide



has no temperature limitations at the operated experimental conditions (Figure 9(d), i.e. this reaction prevails in the high temperature zone, where reactions (1) and (2) are inhibited. Thus, the role of hydrogen in the presented aerosol method can be inferred as preventing the oxidation of the HWG and nanosized catalyst iron particles and also participating in the reaction for the carbon atom release.

35

Example 2. Number distributions of length and diameters of single walled CNTs produced at various conditions and using a ceramic reactor tube.

Carbon source: CO.

5 Catalyst particle source: hot wire generator.

Catalyst material: iron wire of 0.25 mm in diameter.

Operating furnace temperature: 1000, 1200, 1400 °C.

Operating flow rates: hydrogen/nitrogen (7/93) inner flow of 400 cm³/min;

CO outer flow: 400, 590, 765 cm³/min.

10 The example of the CNTs produced at 1200 °C and at equal internal H₂/N₂ and external CO flow rates of 400 cm³/min is described and shown in Example 1.

Number diameter and length distributions of the produced CNTs were obtained on the basis of high-resolution TEM images and presented in Figure 10(a) and Figure 10(b). The investigations of the influence of the experimental conditions
15 on the CNT dimensions were carried out at a fixed hydrogen/nitrogen inner flow of 400 cm³/min varying the furnace temperature from 1000 to 1200 to 1400 °C at a fixed outer CO flow rate of 400 cm³/min and varying the outer CO flow rate from 400 to 590 to 765 cm³/min at a fixed furnace temperature of 1200 °C.

Figure 10(a) shows number length distributions of the produced CNTs. The
20 geometric mean length of CNTs varies from 46 to 54 nm (with the geometric standard deviation between 1.17 and 1.26) with the temperature increase in the system from 1000 to 1400 °C. Increasing the CO flow rate from 400 to 765 cm³/min (or decreasing the residence time) leads to a decrease in the length of CNTs from 54 to 45 nm (with the geometric standard deviation between 1.21 and 1.22).

25 Figure 10(b) shows number diameter distributions of the produced CNTs. The geometric mean diameter of CNTs varies from 0.84 to 1.27 nm (with the geometric standard deviation between 1.24 and 1.40) with the temperature increase in the system from 1000 to 1400 °C. Increasing the CO flow rate from 400 to 765 cm³/min (or decreasing the residence time) leads to a decrease in the length of CNTs
30 from 1.12 to 1.15 nm (with the geometric standard deviation between 1.28 and 1.19).

Figure 10(c) and **Figure 10(d)** show the correlation between diameters of catalyst particles and produced CNTs at different temperatures and residence times (CO flow rates) in the reactor. It can be seen from **Figure 10(c)** that the diameters of CNTs and catalyst particles initiated their growth are correlated and have similar
35 temperature dependence. The length of CNTs can be controlled by the outer CO flow rate, which defines the residence time in the reactor (**Figure 10(d)**). As one can see the dimensions of CNTs such as diameters and lengths can be adjusted by varying the experimental conditions mainly temperature and residence time.

Example 3. Single walled CNT synthesis from carbon monoxide as carbon source using iron as catalyst material and using a stainless steel reactor tube.

Reactor tube: stainless steel with a composition of Fe 53, Ni 20, Cr 25, Mn
5 1.6, Si, C 0.05 weight %.

Carbon source: CO.

Catalyst particle source: hot wire generator.

Catalyst material: iron wire of 0.25 mm in diameter.

Set furnace temperature: 900 °C, corresponding to maximum furnace
10 temperature of around $t_{max} = 1070$ °C.

Operating flow rates: CO outer flow of 400 cm³/min and hydrogen/nitrogen
(7/93) inner flow of 400 cm³/min.

This example, illustrating the synthesis of single walled CNTs, was carried
out in the embodiment of the invention shown in Figure 3(b), wherein the reactor
15 tube was made of stainless steel so as to provide saturated wall conditions for the
iron vapor. Figure 11 demonstrates the product formed at the given operating
conditions. The product consists of bundles of single walled CNTs.

**Example 4. Single walled CNT synthesis from carbon monoxide and
20 octanol/thiophene mixture as carbon sources and reagents and nickel
as catalyst material and using a ceramic reactor tube.**

Carbon source: CO, octanol and thiophene.

Reagent: thiophene (0.5 weight %) and octanol.

Catalyst particle source: hot wire generator.

25 Catalyst material: nickel wire of 0.25 mm in diameter.

Operating furnace temperature: 1200 °C.

Operating flow rates: CO flow of 400 cm³/min and hydrogen/nitrogen (7/93)
flow of 400 cm³/min.

Operating octanol and thiophene vapor pressure in the CNT reactor of 3.4 Pa
30 and 30 Pa.

This example, illustrating the synthesis of single walled CNTs, was carried
out in the embodiment of the invention shown in Figure 3(a). A mixture of
thiophene (0.5 weight %) and octanol was placed in a saturator (6) and was bubbled
at room temperature with carbon monoxide, which was supplied from gas cylinder
35 (1). A ceramic tube, with an internal diameter of 22 mm inserted inside the 40-cm
length furnace (Entech, Sweden) was used as a CNT reactor. Pre-made catalyst
particles were produced in a HWG separated in space from the CNT reactor. The
HWG, which was a resistively heated thin nickel wire, was located inside a glass
bulb. Nickel particles produced by HWG were carried into the CNT reactor with

nitrogen/hydrogen (with mol component ratio of 93.0/7.0) from gas cylinder (2) shown in Figure 3(a). In order to suppress the nickel particle evaporation inside the reactor, the walls of the reactor tube were saturated with nickel by blowing nickel acetylacetonate vapor through the heated up to about 700 °C reactor.

5 In the CNT reactor, the flow carrying the catalyst particles was mixed with the CO flow containing vapors of thiophene and octanol. Inside the CNT reactor, thiophene and octanol decomposition and CO disproportionation took place. It is worth noting that octanol vapor plays two important roles in the CNT reactor: it serves as a carbon source for CNT formation and as a reagent for CNT purification.
10 Formed radicals and fragments containing oxygen after octanol decomposition can easily react with amorphous carbon deposited on the surface of forming CNTs and thus purifies them. Similarly, thiophene was utilized as carbon source and as a reagent. Thiophene supplies sulfur to the catalyst particles. One of the roles of sulfur in the processes of CNT formation is to lower the melting temperature of
15 catalyst particles. Figure 12 demonstrates the product formed at the given operating conditions. CNTs are single walled.

Example 5. Single walled CNT synthesis from ethanol as carbon source and reagent and using iron as catalyst material and using a ceramic reactor tube.
20

Carbon source: ethanol.

Reagent: ethanol.

Catalyst particle source: hot wire generator.

Catalyst material: iron wire of 0.25 mm in diameter.

25 Operating furnace temperature: 1200 °C.

Operating flow rates: hydrogen/nitrogen (7/93) inner flow of 400 cm³/min and nitrogen outer flow of 400 cm³/min.

Operating ethanol vapor pressure in the CNT reactor of 213 Pa.

This example, illustrating the synthesis of single walled CNTs, was carried
30 out in the embodiment of the invention shown in Figure 3(b). Ethanol was placed in a saturator (6) and was bubbled at room temperature with nitrogen, which was supplied from gas cylinder (1). The embodiment consisted of a HWG smoothly integrated with a heated vertical tubular CNT reactor. A ceramic tube, with an internal diameter of 22 mm inserted inside the 90-cm length furnace (Entech,
35 Sweden) was used as a CNT reactor. Nitrogen was supplied from gas cylinder (1). The HWG, which was a resistively heated thin iron wire, was located inside the internal tube. The end of the HWG tube was placed at the location with the CNT reactor wall temperature of about 400 °C. This temperature was found to be optimal,

since reduced particle growth due to the agglomeration and coagulation, minimized particle diffusion losses on the walls and provided a reasonable iron vapor nucleation rate.

5 The metal particles produced by the HWG were carried into the CNT reactor with nitrogen/hydrogen (with mol component ratio of 93.0/7.0) from gas cylinder (2) shown in Figure 3(b). In the CNT reactor, the flow of the metal particles from the HWG was mixed with outer nitrogen flow containing ethanol vapor. Inside the CNT reactor ethanol decomposition took place. It is worth noting that ethanol vapor plays two important roles in the CNT reactor: it serves as a carbon source for CNT
10 formation and as a reagent for CNT purification. Formed radicals and fragments containing oxygen after ethanol decomposition can easily react with amorphous carbon deposited on the surface of forming CNTs and thus purifies them. Figure 13 demonstrates single walled CNT product formed at the given operating conditions. One can see the surface of the produced CNTs does not contain amorphous carbon
15 precipitation and is very clean. Also it is worth noting that virtually all catalyst particles initiated the growth of CNTs. Figure 13 also shows a high-resolution TEM image and the corresponding electron diffraction pattern from a separated SWCNT of 1.6 nm in diameter. One can see from the electron diffraction pattern that the CNT is well crystalline. The radii of the inner and outer circles are consistent with the
20 length of the diffraction vectors of graphite $10\bar{1}0$, and $11\bar{2}0$, respectively. Two sets of spots in the diffraction patterns showing that the CNT is a helical tube.

25 ***Example 6. Single walled and multi-walled CNT synthesis from ethanol/thiophene mixture as carbon sources and reagents and using iron as catalyst material and using a ceramic reactor tube.***

Carbon source: ethanol and thiophene.

Reagent: thiophene (0.5 weight %) and ethanol.

Catalyst particle source: hot wire generator.

Catalyst material: iron wire of 0.25 mm in diameter.

30 Operating furnace temperature: 1200 °C.

Operating flow rates: nitrogen outer flow of 400 cm³/min and inner flow hydrogen/nitrogen (7/93) of 400 cm³/min.

Operating ethanol vapor pressure in the CNT reactor of 2950 Pa and 73 Pa.

Operating thiophene vapor pressure in the CNT reactor of 11 and 0.3 Pa.

35

This example illustrates the possibility to produce both single walled CNTs and multi-walled CNTs depending on the operating conditions, namely, on the vapor pressure of carbon sources (or amount of carbon in the system). A mixture of

thiophene (0.5 weight %) and ethanol was placed in a saturator (6) and was bubbled at room temperature with a carrier gas with and without dilution of flow containing a carbon source. As a result two different ethanol/thiophene vapor pressures in the CNT reactor of 73/0.3 Pa and 2950/11 Pa were obtained. It is worth noting that the smallest operating concentration of carbon sources led to the formation of single walled CNTs, while higher concentration of the alcohol/thiophene mixture led to the production of multi-walled CNTs. Figure 14 and Figure 15 demonstrate the product formed at the given operating conditions and at different ethanol/thiophene vapor pressures. As can be seen from Figure 14, single walled CNTs were produced at the smaller ethanol/thiophene vapor pressures of 73 and 0.28 Pa. Increasing the vapor pressure of the reagent and carbon source (up to 2950 and 11 Pa, respectively) led to the formation of multi-walled CNTs (see Figure 15) and to the formation of amorphous carbon on the surface of the produced CNTs.

Example 7. Fullerene functionalized single walled CNT synthesis from CO as carbon source and hydrogen through a hot wire generator and using iron as catalyst material and using a stainless steel reactor tube.

Reactor tube: stainless steel with a composition of Fe 53, Ni 20, Cr 25, Mn 1.6, Si, C 0.05 weight %.

Carbon source: CO.

Reagent: hydrogen through hot wire generator.

Catalyst particle source: hot wire generator.

Catalyst material: iron wire of 0.25 mm in diameter.

Operating furnace temperature: 900 °C.

Operating flow rates: CO outer flow of 400 cm³/min and inner flow hydrogen of 400 cm³/min.

This example, illustrating the synthesis of fullerene functionalized single walled CNTs, was carried out in the embodiment of the invention shown in Figure 3(b), wherein the reactor tube was made of stainless steel and pure hydrogen was used through the hot wire generator. Figure 16 demonstrates the product formed at the given operating conditions. The product consists of single walled CNTs functionalized fullerene molecules.

Example 8. Fullerene functionalized single walled CNT synthesis from CO as carbon source and hydrogen through hot wire generator and water vapor as a reagent and using iron as catalyst material and using a stainless steel reactor tube.

- 5 Reactor tube: stainless steel with a composition of Fe 53, Ni 20, Cr 25, Mn 1.6, Si, C 0.05 weight %.
- Carbon source: CO.
- Reagent: water vapor at 150 ppm.
- Catalyst particle source: hot wire generator.
- 10 Catalyst material: iron wire of 0.25 mm in diameter.
- Operating furnace temperature: 900 °C.
- Operating flow rates: CO outer flow of 400 cm³/min and inner flow hydrogen/nitrogen (7/93) of 400 cm³/min.
- 15 This example, illustrating the synthesis of fullerene functionalized single walled CNTs, was carried out in the embodiment of the invention shown in Figure 3(b), wherein water vapor was used as a reagent and introduced via a saturator (6) and wherein the reactor tube was made of stainless steel. Figure 17 demonstrates the product formed at the given operating conditions. The product consists of single
- 20 walled CNTs functionalized fullerene molecules.